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Periods and groups valence electrons masses worksheet answers

Which elements have chemical properties similar to those of magnesium? sodium fluorine calcium barium selenium Which elements have chemical properties similar to those of lithium? sodium calcium beryllium barium potassium Which elements have chemical properties similar to those of chlorine? sodium fluorine calcium iodine sulfur Which elements have chemical properties similar to those of carbon? silicon oxygen germanium barium argon Which elements are alkali metals? sodium magnesium aluminum potassium calcium Which elements are alkaline earth metals? sodium magnesium aluminum potassium calcium Which elements are halogens? oxygen fluorine chlorine sulfur carbon Which elements are noble gases? helium hydrogen oxygen neon chlorine Which pairs of elements are located in the same period? H and Li H and He Na and S Na and Rb Which pairs of elements are located in the same period? V and Nb K and Br Na and P Li and Mg In each pair of atoms, which atom has the greater atomic radius? H and Li N and P Cl and Ar Al and Cl In each pair of atoms, which atom has the greater atomic radius? H and He N and F Cl and Br Al and B Scandium is a (metal, nonmetal, semimetal) and is a member of the (main group elements, transition metals). Silicon is a (metal, nonmetal, semimetal) and is a member of the (main group elements, transition metals). 2. 4. 6. 8. 10. 13. metal; transition metals 14. semimetal; main group elements Skills to Develop Give the name and location of specific groups on the periodic table, including alkali metals, alkaline earth metals, noble gases, halogens, and transition metals. Explain the relationship between the chemical behavior of families in the periodic table and their electron configurations. Identify elements that will have the most similar properties to a given element. The chemical behavior of atoms is controlled by their electron configuration. Since the families of elements were organized by their chemical behavior, it is predictable that the individual members of each chemical family will have similar electron configurations. Remember that Mendeleev arranged the periodic table so that elements with the most similar properties were placed in the same group. A group is a vertical column of the periodic table. All of the 1A elements have one valence electron. This is what causes these elements to react in the same ways as the other members of the family. The elements in 1A are all very reactive and form compounds in the same ratios with similar properties with other elements. Because of their similarities in their chemical properties, Mendeleev put these elements into the same group. Group 1A is also known as the alkali metals. Although most metals tend to be very hard, these metals are actually soft and can be easily cut. Group 2A is also called the alkaline earth metals. Once again, because of their similarities in electron configurations, these elements have similar properties to each other. The same pattern is true of other groups on the periodic table. Remember, Mendeleev arranged the table so that elements with the most similar properties were in the same group on the periodic table. It is important to recognize a couple of other important groups on the periodic table by their group name. Group 7A (or 17) elements are also called halogens. This group contains very reactive nonmetal elements. The noble gases are in group 8A. These elements also have similar properties to each other, the most significant property being that they are extremely unreactive, rarely forming compounds. We will learn the reason for this later, when we discuss how compounds form. The elements in this group are also gases at room temperature. An alternate numbering system numbers all of the (s), (p), and (d) block elements from 1-18. In this numbering system, group 1A is group 1; group 2A is group 2; the halogens (7A) are group 17; and the noble gases (8A) are group 18. You will come across periodic tables with both numbering systems. It is important to recognize which numbering system is being used and to be able to find the number of valence electrons in the main block elements regardless of which numbering system is being used. If you can locate an element on the Periodic Table, you can use the element's position to figure out the energy level of the element's valence electrons. A period is a horizontal row of elements on the periodic table. For example, the elements sodium ((\ce{Na})) and magnesium ((\ce{Mg})) are both in period 3. The elements astatine ((\ce{At})) and radon ((\ce{Rn})) are both in period 6. Summary The vertical columns on the periodic table are called groups or families because of their similar chemical behavior. All the members of a family of elements have the same number of valence electrons and similar chemical properties. The horizontal rows on the periodic table are called periods. Vocabulary Group (family): A vertical column in the periodic table. Alkali metals: Group 1A of the periodic table. Alkaline earth metals: Group 2A of the periodic table. Halogens: Group 7A of the periodic table. Noble gases: Group 8A of the periodic table. Transition elements: Groups 3 to 12 of the periodic table. Contributors CK-12 Foundation by Sharon Bewick, Richard Parsons, Therese Forsythe, Shonna Robinson, and Jean Dupon.

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