Charting oxidation number worksheet answers questions pdf printable

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South Pasadena • AP Chemistry

Period ____ Date _/_/_

5 • Reactions in Aqueous Solution

Oxidation Numbers & RedOx

Name_

Example:

Overview:

There are two kinds of reactions in the world, Oxidation-Reduction Reactions and Acid-Base Reactions. In a redox reaction, electrons are gained and electrons are lost.

Double Replacement Reactions are	(redox/non-redox/either)
Single Replacement Reactions are	(redox/non-redox/either)
Synthesis Reactions are	(redox/non-redox/either)
Decomposition Reactions are	(redox/non-redox/either)
Combustion Reactions are	(redox/non-redox/either)

An important idea to help recognize oxidation-reduction reactions it to identify the oxidation number (or oxidation state) of an atom. The oxidation number is also called the "apparent charge." Atoms in ionic compounds have charges and these are also the oxidation numbers. Atoms in molecular compounds can also have oxidation numbers, although they do not have charges.

Rules: The oxidation number of:

- an element in the uncombined state is 0.
- a monatomic ion equals the charge on the ion.
 hydrogen is generally +1; in hydrides, -1.
- oxygen is generally -2; in peroxides, -1.
- oxygen is generally -2; in peroxides, -1.
 elements other than oxygen and hydrogen in a neutral compound is such that the sum of the oxidation numbers for all atoms in the compound is 0.
 elements other than oxygen and hydrogen in a polyatomic ion is such that the sum of the oxidation numbers for all atoms in the ion exclusion hydrogen and hydrogen in the ion
- equals the charge on the ion.

Practice:

1.	Ba ²⁺	2.	<u>N</u> 2	3.	ШI	4.	022·
5.	AgBr	6.	CuCl ₂	7.	HNO3	8.	<u>Cl</u> O ₃
9.	\underline{SO}_3	10.	Na2SO3	11.	BaCrO4	12.	CaSO.
1.	+2	2.	0	3.	+1	4.	-1
5.	-1	6.	+2	7.	+5	8.	+5
9.	+6	10.	+4	11.	+6	12.	+6

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B. +2	16 H		$(90)_{1} \rightarrow O_{1}(90)_{1}$	
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${\rm R}_{\rm c}= {\rm R}_{\rm c}^{2} {\rm S}({\rm R}_{\rm c}^{-}) + {\rm R}_{\rm c}^{-} {\rm R}_{\rm c}^{-$	6	$h_1,h_2 \to h_2$		
Agt Tol + KNO , Sup C. HCBayl + McOHayl -+	3	$=$ $1, P_1 \rightarrow 2P_1$ $\leq -M^2 + 3r_1$		
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Page 1

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Worksheet - ASSIGNING OXIDATION NUMBERS

- Worksheet ASSIGNING OXIDATION NUMBERS
 Oxidation Number Rules:

 The oxidation number of any pure element is 0.
 The oxidation number of a monatomic ion equals that charge on the ion.
 The nore electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.
 The oxidation number of fluorine in a compound is always -1.
 Oxygen has an oxidation number of -2 unless it is combined with F, in which it is +1 or +2, or it is in peroxide (such as H₂O₂ or Na₂O₂), in which it is -1.
 Hydrogen is +1, unless combined with a metal, and then it is -1.
 In compounds, Group 1 is +1, Group 2 is +2, and Aluminum is +3.
 The sum of the oxidation numbers of all atoms in a neutral compound is 0.

Part A: In the following questions, give the oxidation number of the indicated atoms/ion:

1. N in N2O3	16. C in CH4
2. S in H ₂ SO ₄	17. Mn in MnO2
3. C	18. S in SO32.
4. C in CO	19. Mg ²⁺
5. Na in NaCl	20. Cl
6. H in H2O	21. O ₂
7. Ba in BaCl ₂	22. P4
8. N in NO2'	23. Na in Na ₂ S
9. S in Al ₂ S ₃	24. S in H ₂ S
10. S in HSO4	25. Ca ²⁺
11. Cl in Fe(ClO ₂) ₃	26. C in CN
12. Fe in Fe(ClO ₂) ₃	27. H in OH
13. N in NO ₃ :	28. Mn in KMnO ₄



Balancing Redox Reactions Worksheet 1

Balance each redox reaction in acid solution. Mn ²⁺ + BiO₃ \rightarrow MnO₄ $^{-}$ + Bi ³⁺ $\mathsf{MnO4}^{\,\,*} + \,\, \mathsf{S2O3}^{\,\,2^*} \rightarrow \,\, \mathsf{S4O6}^{\,\,2^*} + \,\, \mathsf{Mn}^{\,\,2^+}$ $\text{ClO}_3\ ^{\prime}\ \ast\ \text{Cl}\ ^{\prime}\ \rightarrow\ \text{Cl}_2\ \ast\ \text{ClO}_2$ P + Cu $^{2+} \rightarrow$ Cu + H2PO4 PH3 + 12 → H3PO2 + 1 $NO_2 \rightarrow NO_3$ * NO

Basic Solutions $\mathsf{MnO4}^* \ \ast \ \mathsf{C2O4}^{\ 2^*} \ \rightarrow \mathsf{MnO2} \ \ast \ \mathsf{CO2}$

CIO2 → CIO2 + CIO3 *

 $\text{Cu(NH3)4}^{2+} \ \ast \ \text{S2O4}^{2^{*}} \rightarrow \ \text{SO3}^{2^{*}} + \ \text{Cu} \ \ast \ \text{NH3}$

Zn + NO3 $^{\circ}$ \rightarrow Zn(OH)4 2* + NH3

AI * OH ' \rightarrow AIO2 ' * H2

 $Zn \rightarrow Zn(OH)_4 \overset{2-}{} + H_2$

Chemistry 30 Unit 7: Redox Reactions and Electrochemistry Ans Practice Assignment 7.1 - Oxidation Numbers and Redox Reactions

Answers

1. Determine the oxidation number of each element in the following compounds.

Rules: 1. Pure elements have an oxidation number of 0

- If the compound is an ionic compound, the oxidation number for each element is the ion's charge
- The oxidation number of hydrogen in a compound is +1
- The oxidation number of oxygen in most compounds is -2
- (peroxides are the exception; in peroxides oxygen has an oxidation number of -1)
- The sum of the oxidation numbers in a compound is zero.
- The sum of the oxidation numbers in a polyatomic ion is equal to the ion charge.

		Hint	Oxidation Numbers for each Element				
а.	SnCli	Rule 2	Sn	+4	CI	-1	
b.	Ca ₃ P ₂	Rule 2	Ca	+2	Р	-3	
C.	SnO	Rules 4, 5	Sn	+2	0	-2	
d.	Ag ₂ S	Rule 2	Ag	+1	s	-2	
e.	н	Rule 3, 5	н	+1	Ĩ,	-1	
f.	N2H4	Rule 3, 5	Ν	-2	н	+1	watch the sign for N!
9.	Al ₂ O ₃	Rule 4, 5	AI	+3	0	-2	
h.	Sa	Rule 1	s	0			
i,	HNO ₂	Rules 3, 4, 5	н	+1	N	+3	0 <u>-2</u>
j.	O ₂	Rule 1	0	0			pure element!
k.	H3O*	Rules 3, 4, 6	н	+1	0	-2	surprised?
Ι.	CIO3.	Rules 4, 6	CI	+5	0	-2	
m.	S ₂ O ₃ ^{2.}	Rules 4, 6	s	+2	0	-2	
n.	KMnO ₄	Rules 4, 5, 6	к	+1	Mn	+7	02
0.	(NH4)2SO4	Rules 4, 5, 6	N	-3	н	+1	S <u>+6</u> O <u>-2</u>

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