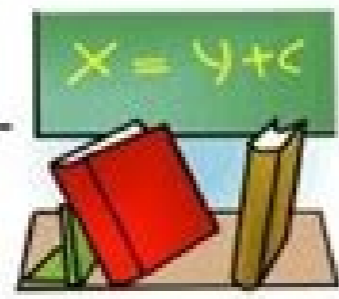


I'm not robot!

Name: _____

Basic Algebra



Evaluate each expression.

$a = 3, \quad b = 5, \quad c = 6$

- 1. $a + 5$ _____
- 2. $15 - c$ _____
- 3. $4b$ _____
- 4. $\frac{18}{c}$ _____
- 5. $20 - a$ _____
- 6. $11b$ _____
- 7. $\frac{45}{b}$ _____
- 8. $a - 2$ _____
- 9. $a + b + c$ _____
- 10. $\frac{c}{a}$ _____

$p = 12, \quad q = 2, \quad r = 30$

- 11. $q50$ _____
- 12. $\frac{r}{q}$ _____
- 13. $p + 4 + 6$ _____
- 14. $p - 7$ _____
- 15. $10r$ _____
- 16. $\frac{r}{10}$ _____
- 17. $\frac{p}{4}$ _____
- 18. $r - p$ _____

7th Grade Holt Mathematics
Quiz 2-4 to 2-6

Name: _____

Find the prime factorization (Use EXPONENTS if necessary). Each worth 2 pts

- _____ 1. 56
- _____ 2. 480

Find the GCF for: Each worth 2 pts:

- _____ 3. 28 and 98
- _____ 4. 21 and 24.

Find the LCM for: Each worth 2 pts

- _____ 5. 12 and 90
- _____ 6. 8, 12 and 30

Write down if the following numbers are composite or prime: Each worth 1 pt

- _____ 7. 11
- _____ 8. 15
- _____ 9. 18
- _____ 10. 29



Adding mixed numbers and fractions (unlike denominators)

Grade 5 Fractions Worksheet

Find the sum.

- 1. $5\frac{2}{3} + \frac{5}{7} =$ _____
- 2. $7\frac{3}{8} + \frac{1}{12} =$ _____
- 3. $1\frac{1}{10} + \frac{7}{8} =$ _____
- 4. $5\frac{4}{7} + \frac{1}{9} =$ _____
- 5. $2\frac{3}{5} + \frac{1}{5} =$ _____
- 6. $1\frac{8}{11} + \frac{2}{8} =$ _____
- 7. $10\frac{6}{10} + \frac{2}{4} =$ _____
- 8. $2\frac{1}{2} + \frac{5}{8} =$ _____
- 9. $6\frac{6}{7} + \frac{3}{8} =$ _____
- 10. $10\frac{2}{8} + \frac{5}{12} =$ _____
- 11. $1\frac{4}{11} + \frac{2}{5} =$ _____
- 12. $7\frac{8}{11} + \frac{3}{5} =$ _____
- 13. $7\frac{3}{10} + \frac{5}{12} =$ _____
- 14. $7\frac{2}{5} + \frac{1}{4} =$ _____

Elements, Compounds, & Mixtures Practice

Part 1 - Circle the elements in the list below:

<u>Iron</u>	Carbon dioxide	wood	iron	<u>oxygen</u>
water	<u>hydrogen</u>	<u>nitrogen</u>	<u>hydrogen</u>	<u>oxygen</u>
<u>Iron</u>	light	fat	air	<u>nitrogen</u>
<u>hydrogen</u>	<u>carbon</u>	<u>oxygen</u>	<u>iron</u>	<u>oxygen</u>

Part 2 - Classify the following as a pure substance (PS) or mixture (M).

air <u>M</u>	gasoline <u>M</u>	water (H ₂ O) <u>PS</u>	seawater (H ₂ O) <u>PS</u>	gold (Au) <u>PS</u>
crushed ice <u>PS</u>	orange (C) <u>PS</u>	sea water <u>M</u>	pure <u>M</u>	acid <u>M</u>

Part 3 - Classify the following as an element (E), compound (C), solution (S), or a heterogeneous mixture (HM).

hydrogen <u>E</u>	salt water <u>HM</u>	carbon dioxide (CO ₂) <u>C</u>	water (H ₂ O) <u>C</u>	orange juice <u>S</u>
air (C) <u>E</u>	table salt (NaCl) <u>C</u>	mercury (Hg) <u>E</u>	air (mixture) <u>HM</u>	periodic table <u>HM</u>
aluminum (Al) <u>E</u>	air <u>HM</u>	gasoline <u>S</u>	water and instant coffee <u>S</u>	
air <u>HM</u>	steam (H ₂ O) <u>C</u>	table salt (NaCl) <u>C</u>	orange juice <u>S</u>	periodic table <u>HM</u>

Part 4 - Classify the following as an element (E), compound (C), or mixture (M). If it does not fit under any of these categories, write an X.

<u>E</u> diamond (C)	<u>C</u> sugar (C ₆ H ₁₂ O ₆)	<u>M</u> air	<u>E</u> iron (Fe)
<u>M</u> air	<u>C</u> sulfuric acid (H ₂ SO ₄)	<u>M</u> gasoline	<u>S</u> seawater
<u>E</u> oxygen (O)	<u>E</u> boron (B)	<u>E</u> uranium (U)	<u>M</u> pudding
<u>C</u> water (H ₂ O)	<u>C</u> glucose (C ₆ H ₁₂ O ₆)	<u>M</u> instant noodle	<u>M</u> a dog
<u>C</u> ammonia (NH ₃)	<u>C</u> salt (NaCl)	<u>S</u> orange	<u>C</u> gold (Au)
<u>M</u> wood	<u>M</u> bronze	<u>M</u> air	<u>M</u> pizza
<u>C</u> dry ice (CO ₂)	<u>C</u> baking soda (NaHCO ₃)	<u>E</u> uranium (U)	<u>M</u> orange

Part 5 - Classify each of the pictures below as an Element, Mixture, Compound, Mixture of Elements, Mixture of Compounds, or Mixture of Elements and Compounds.

<u>compound</u>	<u>mixture of elements</u>	<u>mixture of compounds</u>
<u>element</u>	<u>mixture of elements</u>	<u>mixture of compounds</u>

A Sure-Fire Way to Draw Lewis Structures!

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Lewis structures are a way to write chemical compounds where all the atoms and electrons are shown. Sometimes, people have a lot of trouble learning how to do this. However, using the methods on this page, you should have very little trouble.

The first method given allows you to draw Lewis structures for molecules with no charged atoms, while the second allows you to do it for charged molecules (such as polyatomic ions).

How to draw Lewis structures for molecules that contain no charged atoms

- 1) **Count the total valence electrons for the molecule:** To do this, find the number of valence electrons for each atom in the molecule, and add them up.
- 2) **Figure out how many octet electrons the molecule should have, using the octet rule:** The octet rule tells us that all atoms want eight valence electrons (except for hydrogen, which wants only two), so they can be like the nearest noble gas. Our first goal is to figure out how many electrons each atom in the molecule should have, and add them up. The only weird element is boron - it wants six electrons.
- 3) **Subtract the valence electrons from octet electrons:** Or, in other words, subtract the number you found in #1 above from the number you found in #2 above. The answer you get will be equal to the number of bonding electrons in the molecule.
- 4) **Divide the number of bonding electrons by two:** Remember, because every bond has two electrons, the number of bonds in the molecule will be equal to the number of bonding electrons divided by two.
- 5) **Draw an arrangement of the atoms for the molecule that contains the number of bonds you found in #4 above:** Some handy rules to remember are these:
 - Hydrogen and the halogens bond once.
 - The family oxygen is in bonds two times.
 - The family nitrogen is in bonds three times. So does boron.
 - The family carbon is in bonds four times.
- 6) **Find the number of lone pair (nonbonding) electrons by subtracting the bonding electrons (#3 above) from the valence electrons (#1 above). Arrange these around the atoms until all of them satisfy the octet rule:** Remember, ALL elements EXCEPT hydrogen

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